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Circle the  
letter of the  
best answer. 1.

Given the  
balanced

equation  $2\text{C}_4\text{H}_{10} + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 10\text{H}_2\text{O}$ , how  
many moles of  $\text{CO}_2$   
are produced  
when 14.9 g of  $\text{O}_2$   
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1. 88% The  
actual yield of  
a reaction is 22  
g and the  
theoretical  
yield is 25 g.

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mixed with 12.0

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SECTION 3 SHORT

ANSWER Answer

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Chemistry in the  
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- Chapter 9  
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1. a The notation for sodium chloride,  $\text{NaCl}$ , stands for one (a) formula unit. (c) crystal. (b) molecule. (d) atom. 2. d In a crystal of an ionic compound, each cation is

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surrounded by a  
number of (a)  
molecules. (c)  
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