

Chapter 6 Chemical Bonds Answers

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[Chapter 6 Chemical Bonds WordWise](#). STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. JolieDA. Key Concepts: Terms in this set (12) ionic. A type of bond that holds cations and anions together. Covalent. A type of bond in which two atoms share a pair of valence electrons. Molecule.

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[CHAPTER 6 REVIEW Chemical Bonding SECTION 3 SHORT ANSWER](#) Answer the following questions in the space provided. 1. a The notation for sodium chloride, NaCl, stands for one (a) formula unit. (c) crystal. (b) molecule. (d) atom. 2. d In a crystal of an ionic compound, each cation is surrounded by a number of (a) molecules. (c) dipoles. (b) positive ions. (d) negative ions.

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Section 6.2 – Covalent Bonding A covalent bond is a chemical bond in which two atoms share a pair of valence electrons. When two atoms share one pair of electrons, the bond is called a single bond.

Covalent vs Ionic Bond

~~Chapter 6 – Chemical Bonds~~

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Oxygen has higher electronegativity than sulfur, which creates a highly polar bond. Increase polarity in H₂O bonds means a stronger intermolecular attraction making water a liquid at room temperature. Hydrogen bonding exists between water molecules, but not between hydrogen sulfide molecules.

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Play this game to review Atoms & Molecules. How many electrons should Helium have around its Lewis dot model?

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Ch. 2 Answer Key - lawndalehs.org. Section Review 2-1 1. protons; neutrons 2. electrons 3. neutrons 4. electrons 5. ionic 6. The two main types of chemical bonding are ionic and covalent bonding. Ionic bonds are formed when a transfer of electrons takes. <http://www.lawndalehs.org/ourpages/auto/2006/9/12/1158089916202/Ch%202%20answer%20key.pdf...>

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- A covalent bond is a chemical bond in which two atoms share a pair of valence electrons.
- When two atoms share one pair of electrons, the bond is called a single bond. When two atoms share two pairs of electrons, the bond is called a double bond.
- A molecule is a neutral group of atoms that are joined together by one or more covalent bonds.

~~Chapter 6 Chemical Bonds – Henry County School District~~

Chapter 6: Chemical Bonds Review Answer Key. What is the most stable group on the Periodic Table? Why is this group stable? Noble Gases (Group 18) because they have 8 valence electrons which means their outer energy level is full. Name the most reactive group(s) on the Periodic Table? Why are they reactive?

Name _____

each statement or best answers each question. 1 If two covalently bonded atoms move closer than a distance of the bond length, the potential energy of the atoms a. becomes negative. b. decreases. increases. d. remains constant. 2. The electrons involved in the formation of a covalent bond are a. transferred from one atom to another.

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Modern Chemistry 13 Chemical Bonding CHAPTER 6 STUDY GUIDE Chemical Bonding SECTION 5 MOLECULAR GEOMETRY SHORT ANSWER Answer the following questions in the

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space provided. 1. Identify the major assumption of the VSEPR theory, which is used to predict the shape of atoms.

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