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increase i.... the gain of electrons, which also corresponds to a decrease in.... The flow of electric charge. A device in which a chemical reaction either produces or is ca....

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Answer:  $E_{\text{cell}} = -0.22 \text{ V}$ ; the reaction will not occur spontaneously. Applying the Nernst equation to a simple electrochemical cell such as the Zn/Cu cell discussed in Section 19.2 allows us to see how the cell voltage varies as the reaction progresses and the concentrations of the dissolved ions change.

Chapter 19.4: Electrochemical Cells and Thermodynamics ...

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Chapter 19: Electrochemistry Dr. Pahlavan / Dr. Ghanbaripour 3 9. Given the  
following data:  $2\text{Fe}^{2+} + 2\text{e}^{-} \rightarrow \text{Fe}$   $E_{\text{red}} = -0.44 \text{ v}$ ,  $\text{Ag}^{+} + \text{e}^{-} \rightarrow \text{Ag}$   $E_{\text{red}} = 0.80 \text{ v}$   
Answer the following questions with respect to the reaction;  $\text{Fe}^{2+}(\text{aq}) + 2 \text{Ag}(\text{s}) \rightarrow$   
 $\text{Fe}(\text{s}) + 2 \text{Ag}^{+}(\text{aq})$

Chapter 19: Electrochemistry Dr. Pahlavan / Dr ...

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reduction. electric current. electrochemical charge. the loss of electrons, which also  
corresponds to an increase i.... the gain of electrons, which also corresponds to a  
decrease in.... The flow of electric charge. A device in which a chemical reaction

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The hydrogen-oxygen fuel cell is described in Section 19.6 (a) What volume of  
 $\text{H}_2(\text{g})$ , stored at  $25^{\circ} \text{C}$  at a pressure of 155

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atm, would be needed to run an electric motor drawing a current of 8.5 A for \$3.0  
 (b) What volume (in liters) of air at  $25^{\circ}\text{C}$  and 1.00 atm will have to pass into the cell per minute to run the motor?

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Chapter 19 Electrochemistry Chang 4 19.4 SPONTANEITY OF REDOX REACTIONS

Electrical work is needed to move a charge through a conductor: Electrical energy = charge x potential difference Units: Joules = Coulombs x Volts F is Faraday's constant; the electrical charge contained in 1 mol of e-is equal to 96500 C. 1 F = mole– Coulombs 1 96500 = V mole–

### CHAPTER 19 ELECTROCHEMISTRY

Read Chapter 18: Entropy, Free Energy, and Equilibrium & Read Chapter 19:

Electrochemistry Answer the following problems in the space provided. For problems involving an equation, carry out the following steps: 1. Write the equation. 2. Substitute numbers and units. 3. Show the final answer with units. There is no credit without showing work.

AP Chemistry Chapter 18 & 19: Thermodynamics ...

Chapter 19 Electrochemistry Math Summary Relating Standard Cell Potential to Standard Half Cell Potentials  $E^{\circ}_{\text{cell}} = E^{\circ}_{\text{oxidation}} + E^{\circ}_{\text{reduction}}$  (standard conditions assume 1.0 M concentrations) Relating Half Cell Potentials when Written in

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Opposite Directions  $E^{\circ}_{\text{ox}} = -E^{\circ}_{\text{red}}$  for half reactions written in opposite directions

Chapter 19 Electrochemistry Math Summary

Answer:  $\text{MnO}_4^- (\text{aq}) + 8\text{H}^+ (\text{aq}) + 5\text{e}^- \rightarrow \text{Mn}^{2+} (\text{aq}) + 4\text{H}_2\text{O} (\text{l})$ ;  $\text{Sn}^{2+} (\text{aq}) \rightarrow \text{Sn}^{4+} (\text{aq}) + 2\text{e}^-$  The Pt electrode in the permanganate solution is the cathode; the one in the tin solution is the anode.

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