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IB Chemistry HL: Unit 14: Acids & Bases ANSWERS 4/10 Short Answers 1)

(a) In the reaction  $2\text{H}_2\text{O}(l) \rightleftharpoons \text{H}_3\text{O}^+(aq) + \text{OH}^-(aq)$  use the Bronsted-Lowry Theory to discuss the acidic and/or basic nature of water. [2]  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  Water can donate a proton (BL acid) or receive a proton (BL base)

## **ANSWERS Unit 14: Review Acids and Bases**

acids react with bases in a neutralization reaction( also double replacement). applications: the reaction of an acid and a base produces a salt and water. List the common strong acids and the strong bases.

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Both can also be described as having unique tastes and capable of changing the color of indicators. Describe the process of dissociation of phosphoric acid ( $H_3PO_4$ ). During the dissociation of acid, the solute ( $H_3PO_4$ ) is dissolved in water.

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the subject.

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Acid and Base Worksheet - Answers. 1) Using your knowledge of the Brønsted-Lowry theory of acids and bases, write equations for the following acid-base reactions and indicate each conjugate acid-base pair: a)  $\text{HNO}_3 + \text{OH}^- \rightarrow \text{H}_2\text{O} + \text{NO}_3^-$ .  $\text{HNO}_3$  and  $\text{NO}_3^-$  make one pair  $\text{OH}^-$

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and H<sub>2</sub>O make the other. b)  $\text{CH}_3\text{NH}_2 + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{NH}_3^+ + \text{OH}^-$

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A Base is any substance whose aqueous solution is characterized by a bitter taste, a slippery feel, the ability to turn red litmus blue, and the ability to react with acids to form salts. Questions...

## **Answers about Acids and Bases**

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Solution: The reaction of the acid + base gives a product of salt + water, which is considered as neutralization reaction. Examples:  $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

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indicate the conjugate acid-base pairs. base 2  $\text{CO}_3^{2-}$  acid 2  $\text{HCO}_3^-$   
base 1  $\text{NH}_3$  True b. A proton has been transferred from acid 1 to base 2 in the above reaction. True or False? 3. Consider the neutralization reaction described by the equation:  $\text{HCO}_3^-(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{CO}_3^{2-}(\text{aq}) + \text{H}_2\text{O}(\text{l})$  a. Label the conjugate acid-base pairs in this system.  $\text{HCO}_3^-$

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acid - substances that ionize in solutions to form  $H^+$  ions.  
amphoteric - a substance that can be an acid or a base. Arrhenius Model - in aqueous solutions, acids form hydrogen ions ( $H^+$ ). base - substances that ionize in solutions and form  $OH^-$  ions. binary acids - acids that do not contain oxygen in their chemical formula.

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